With these corrections our structure determination is in essential agreement with that of Davies, Mais, and Owston, *J. Chem. Soc., A*, 1750 (1967).—E. L. AMMA

D. M. Tully-Smith, R. K. Kurimoto, D. A. House, and C. S. Garner: Hydrolysis Kinetics of the Blue Isomer of Dichlorodiaquoethylenediaminechromium(III) Cation and of the Magenta Isomer of Chlorotriaquoethylenediaminechromium(III) Cation.

Page 1527. In column 2, paragraph 3, we reported that the contribution of base hydrolysis, if any, was unknown in the hydrolysis of magenta $Cr(en)(OH_2)_3Cl^{2+}$. We now have obtained k_2 in 0.35 and 0.07 F $HClO_4$ (μ = 1.5 M, $NaClO_4$) at 60°; the rates at these two acid concentrations, together with the earlier value in 1.5 F $HClO_4$, give a linear plot of k_2 *vs.* $1/[H^+]$. If we interpret the $[H^+]$ dependence in terms of contribution from aquation of $Cr(en)(OH_2)_2(OH)Cl^+$, which appears to be the only reasonable explanation, we obtain $k_{2a} = 3.2 \times 10^{-5}$ sec^{-1} at 60° for aquation of magenta $Cr(en)(OH_2)_3Cl^{2+}$ and (estimating the first acid dissociation constant of the chlorotriaquo ion as $K_1 \sim 10^{-4}$ at 60°) $k_{2h} \sim 0.03$ sec^{-1} at 60° for aquation of the hydroxodiaquo complex. Base hydrolysis is considered to contribute negligibly.

Page 1527. In column 2, lines 7–9 of the fourth paragraph, the sentence starting "This value . . ." should read: This value is *ca.* one-hundredth of the Cl^--release hydrolysis rate constant of *cis*-$Cr(en)_2(OH_2)Cl^{2+}$ in 0.1 F HNO_3,[7] and about one-tenth the constant for *trans*-$Cr(en)_2(OH_2)Cl^{2+}$ Cl^- release . . .

These corrections are in addition to those noted earlier in *Inorg. Chem.*, **6**, 2273 (1967).—C. S. GARNER

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R. Graham Hughes and Clifford S. Garner: The Use of Chromium(IV) Diperoxo Amines in the Synthesis of Chromium(III) Amine Complexes. IV. Some Chloro and Bromo Ammineethylenediamine Complexes.

Page 76. In column 1, line 17, isomers I and II were accidentally interchanged; isomer I presumably has all three H_2O ligands *cis* to one another.—C. S. GARNER

John M. Malin and James H. Swinehart: The Kinetic and Equilibrium Properties of the Vanadium(II) Thiocyanate Complex in Aqueous Solution.

Page 250. In line 5 of the abstract and subsequent places in the text where it appears (page 253; column 1, line 3 and column 2, line 17) the entropy of activation for the bimolecular rate constant for complex formation should be -7 ± 4 eu and not -2 ± 4 eu—J. H. SWINEHART.

Colin F. Bell and David R. Rose: Complexes between Pyridine-2-aldehyde-2′-pyridylhydrazone and the Platinum Metals. I. Rhodium.

Page 325. In the above paper, we used the nomenclature of R. Kuhn and W. Münzing for the geometrical isomers of hydrazones. We also referred to the work of Karabatsos, who uses the opposite nomenclature. Nomenclature is consistent throughout our paper if the following changes are made in column 2 of page 328: line 19, for *syn* read *anti;* line 20, for *anti* read *syn;* line 22, for *syn* read *anti*. Definitions are according to R. Kuhn and W. Münzing, *Chem. Ber.*, **85**, 29 (1952).—C. F. BELL.

Ruth A. Goodrich and P. M. Treichel: Synthesis and Characterization of the Alkylhydridotrifluorophosphoranes CH_3PF_3H and $C_2H_5PF_3H$.

Page 696. In Table II, the proton chemical shift values should be negative; *i.e.*, for CH_3PF_3H, $H_{P-H} = -7.1$, $H_{CH_3} = -1.7$; for $C_2H_5PF_3H$, $H_{P-H} = -7.4$, $H_{CH_3} = -1.5$, $H_{CH_2} = -0.5$. Footnote *a* should read "Relative to $CFCl_3$."—P. M. TREICHEL

R. F. Childers, Jr., K. G. Vander Zyl, Jr., D. A. House, R. G. Hughes, and C. S. Garner: Synthesis of a Monodentate Ethylenediamine Complex of Chromium(III) and Kinetics of Hydrolysis of Tetraaquoethylenediaminechromium(III) Cation and of Pentaaquo(2-aminoethylammonium)chromium(III) Cation.

Page 749. In line 2 of abstract, and in column 2, paragraph 2, we reported that so far as we knew, we were reporting the first isolation of a complex of chromium(III) with a monodentate ethylenediamine ligand. Professor Gerold Schwarzenbach has kindly informed us that isolation of the red chloride salt of μ-hydroxo-enneaammine(2-aminoethylammonium)dichromium(III) cation, $[(H_3N)_5Cr(OH)Cr(enH)(NH_3)_4]Cl_6$, which is a chromium(III) complex with a monodentate ethylenediamine ligand, was reported earlier [G. Schwarzenbach and B. Magyar, *Helv. Chim. Acta*, **45**, 1454 (1962)]. Synthesis was achieved by treatment of $(H_3N)_5Cr(OH)Cr(NH_3)^{5+}$ with anhydrous ethylenediamine, giving the blue "basic rhodo" cation, $(H_3N)_5CrOCr(en)(NH_3)_4^{5+}$, which was converted in solution by hydrochloric acid to the μ-hydroxo compound.

Page 753. In column 2, paragraph 2, line 3, the first chemical formula should read: $Cr(en)(OH_2)_4^{3+}$.—C. S. GARNER

R. Graham Hughes and Clifford S. Garner: Aquation Kinetics of the Green-Blue Isomer of Dibromodiaquoethylenediaminechromium(III) Cation and the Magenta Isomer of Bromotriaquoethylendiaminechromium(III) Cation.

Page 1989. In column 1, the chemical formula in the heading under Results and in eq 1 should be $Cr(en)(OH_2)_2Br_2^+$. In column 2, footnote *d* of Table I, "deviations" should be in the singular.

Page 1991. In Table III, the units of ΔS^* should be cal deg^{-1} mol^{-1} and the sixth complex in the first column should be *cis*-$Cr(en)_2Br_2^{+\,c}$.

Page 1992. In column 1, line 2, the chemical formula should be $Cr(en)(OH_2)_3Br^{2+}$.—C. S. GARNER